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11.3 Reactions in Aqueous Solution 28 >  
You have seen that mixing solutions of two ionic compounds can sometimes result in the formation of an insoluble salt called a precipitate. •Some combinations of solutions produce precipitates, while others do not.

## **11.3 Reactions in Aqueous Solution - Quia**



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The reaction of aqueous solutions of silver nitrate with sodium chloride to form solid silver chloride and aqueous sodium nitrate is a double-replacement reaction. The reaction is shown in Figure 11.11.  $\text{AgNO}_3(\text{aq}) + \text{NaCl}(\text{aq}) \rightarrow \text{AgCl}(\text{s}) + \text{NaNO}_3(\text{aq})$  This is the way you have been writing equations involving aqueous solutions of ionic compounds.

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AP Chem - Chapter 4:  
Aqueous Reactions and  
Solution Stoichiometry.

STUDY. PLAY. aqueous  
solutions (p.113)

solutions in which  
water is the dissolving  
medium. precipitation

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reactions (p.113)  
soluble reactants yield  
an insoluble product.  
acid-base reactions  
(p.113)  $H^+$  ions are  
transferred between  
reactants.

## **AP Chem - Chapter 4: Aqueous Reactions and Solution ...**

Assume that an  
aqueous solution of  
hydroxide  
ion,  $OH^-$ , represented by  
unshaded spheres, is

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allowed to mix with a solution of an acid,  $H_nA$ , represented by gray spheres. Three possible outcomes are represented by boxes (a)-(c), where the black spheres represent  $A^{n-}$ , the anion of the acid. For clarity  $H_2O$  molecules are not shown.

## **Quiz+ | Quiz 4: Reactions in Aqueous Solutions**

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an Inverse Electron  
Demand Hetero  
Diels–Alder Reaction in  
Aqueous Solution and  
Organic Solvents:  
Cycloaddition of  
Substituted Styrenes to  
Di(2-pyridyl)-1,2,4,5-tet  
razine. The Journal of  
Organic Chemistry  
1996 , 61 (6) ,  
2001-2005.

**Diels-Alder reactions  
in aqueous  
solutions. Enforced**

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Abstract Chemical reactions in Aqueous solutions In this lab we are focusing on combining two substances to make a chemical reaction. The purpose of this lab is to have an understanding of the concept of insoluble and soluble, as well as oxidized and reduced agents. By the



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end of the lab, we will be able to understand how to make a net ionic equation based on the observations and the ...

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