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A balanced chemical equation allows

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one to determine A. The mole ratio of any two substances in the reaction B. The energy released in the reaction C. The electron configuration of all elements in the reaction D. The mechanism involved in the reaction

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Identify the letter of the choice that best
completes the statement or answers the
question. 1. If you start with 10.0 moles
of C₃H₈(propane) and 10.0 moles of O₂,
what is the limiting reactant. C₃H₈(g) +
5O₂(g) → 3CO₂(g) + 4H₂O(g) a.

Chapter 9 Prep-Test

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Chapter 9 Practice Test Multiple Choice
Identify the letter of the choice that best completes the statement or answers the question. ____ 1. A solid produced by a chemical reaction in solution that separates from the solution is called a. a precipitate. c. a molecule. b. a reactant.

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296 Chapter 9 • Chemical Reactions.
Figure 9.15 When aqueous sodium hydroxide is added to a solution of copper(II) chloride, the anions (OH^- and Cl^-) change places. The resulting products are sodium chloride, which remains in solution, and copper(II) hydroxide, the blue solid in the test tube.

Chapter 9: Chemical Reactions

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