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Koretsky helps students understand and visualize thermodynamics through a qualitative discussion of the role of molecular interactions and a highly visual presentation of the material. By showing how principles of thermodynamics relate to molecular concepts learned in prior courses, *Engineering and Chemical Thermodynamics, 2e* helps students construct new knowledge on a solid conceptual foundation.

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Milo D. Koretsky received his Ph.D. in Chemical Engineering from the University of California at Berkeley in 1991. He is currently of professor of Chemical Engineering at Oregon State University.

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Koretsky's qualitative discussion of the role of molecular interactions and the visual approaches he uses helps students understand and visualize thermodynamics.

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Overview. Chemical engineers face the challenge of learning the difficult concept and application of entropy and the 2nd Law of Thermodynamics. By following a visual approach and offering qualitative discussions of the role of molecular interactions, Koretsky helps them understand and visualize thermodynamics.

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I received my B.S. in chemical engineering (2005) at OSU and studied this text, in part, under Dr. Koretsky. The text was split into two undergraduate courses. I recently entered a PhD program at Texas A&M and reviewed almost the entire text in preparation for the qualifier.

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Chapter 1 Solutions Engineering and Chemical Thermodynamics
Wyatt Tenhaeff Milo Koretsky Department of Chemical
Engineering Oregon State University koretsm@enr.orst.edu f1.2
An approximate solution can be found if we combine Equations
1.4 and 1.5: $1 \text{ } \square 2 \text{ mV } \square \text{ ekmolecular } 2 \text{ } 3 \text{ kT } \square \text{ ekmolecular } 2 \text{ } \square 3\text{kT}$
 $\square \text{V } \square \text{ m}$ Assume the temperature is $22 \text{ } ^\circ\text{C}$.

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