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greater than 1.7 and is
a transfer of electrons.
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transfers to the higher
electronegativity) t.
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a crystal lattice
structure. Lattice
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CHAPTER 6 REVIEW

Chemical Bonding

SECTION 4 SHORT

ANSWER Answer the
following questions in

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the space provided. 1.
b In metals, the
valence electrons are
considered to be (a)
attached to particular
positive ions. (c)
immobile.

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Chemical Bonding 1
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bond- a mutual electrical attraction between the nuclei and the valence electrons of different atoms that binds the atoms together ; 2. ionic bond- a chemical bond that results from the electrical attraction between anions and cations ...

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6. Metals are malleable

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because when struck, one plane of metal atoms can slide past another plane without breaking bonds. b. cannot easily move out of the way. c. moves in a way that maximizes the repulsive forces within the

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titrated with a weak base, the pH of the solution at the equivalence point is than 7. 19. A is a highly purified solid used to check the concentration of a standard solution. 20. A 1 M solution of NaOH will have a pH that is

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ANSWER Answer the following questions in the space provided. 1.
a.

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Answer Key
Answer to 12. Each period in the periodic table corresponds to .
a. an energy level c. an orbital
b. an electron d. a suborbital

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