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In chemistry,
colligative properties
are those properties of
solutions that depend
on the ratio of the
number of solute
particles to the number

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of solvent particles in a solution, and not on the nature of the chemical species present. The number ratio can be related to the various units for concentration of a solution, for example, molarity, molality, normality (chemistry), etc.

Colligative properties - Wikipedia

When this solution is

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actually prepared and its freezing point depression measured, however, a value of 3.4°C is obtained. Similar discrepancies are observed for other ionic compounds, and the differences between the measured and expected colligative property values typically become more significant as solute concentrations increase.

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11.4 Colligative Properties - Chemistry

A colligative property is a property of a solution that is dependent on the ratio between the total number of solute particles (in the solution) to the total number of solvent particles. Colligative properties are not dependent on the chemical nature of the solution's components.

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Colligative Related **Properties - Definition, Types, Examples ...**

The colour of a solution is another non-colligative characteristic. In contrast to the colourless salt and sugar solutions, a 0.5 M solution of CuSO_4 is vivid blue. Viscosity, surface tension, and solubility are examples of non-colligative

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characteristics.

Colligative Properties - Definition, Facts & Mathematical ...

Freezing point depression is a colligative property observed in solutions that results from the introduction of solute molecules to a solvent. The freezing points of solutions are all lower than that of the pure solvent and is directly

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proportional to the molality of the solute.

$$\Delta T_f = T_f(\text{solvent}) - T_f(\text{solution}) = K_f \times m$$

Freezing Point Depression - Chemistry LibreTexts

Solution: Option (i) is the answer. 8.

Colligative properties depend on _____. (i) the nature of the solute particles dissolved in

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solution. (ii) the number of solute particles in solution.

(iii) the physical properties of the solute particles dissolved in solution. (iv) the nature of solvent particles.

Solution: Option (ii) is the answer. 9.

I. Multiple Choice Questions (Type-I) 1. Which of the ...

COLLIGATIVE
PROPERTIES Colligative properties depend only

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on the number of solute particles present, not on the identity of the solute particles. Among colligative properties are Vapor pressure lowering Boiling point elevation Melting point depression Osmotic pressure

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strategy here will be to determine the van't Hoff factor for sodium iodide, "NaI" calculate the molality of the solution calculate the freezing-point depression of the solution The idea is that the freezing point of a solution is lower than the freezing point of the pure solvent, which for water is 0°C at normal pressure. Now, the difference between the

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freezing point of the ...

Property Related

What is the freezing point of a solution that contains 0.5 ...

Vapour pressure, pressure exerted by a vapour when the vapour is in equilibrium with the liquid or solid form, or both, of the same substance—i.e., when conditions are such that the substance can exist in both or in all three phases. Vapour

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pressure is a measure of the tendency of a material to change into the gaseous or vapour state, and it increases with temperature.

vapour pressure | Definition & Facts | Britannica

The freezing point depression due to the presence of a solute is also a colligative property. That is, the amount of change in the freezing point is

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related to the number of particles of solute in a solution and is not related to the chemical composition of the solute.

13.9: Freezing Point Depression and Boiling Point ...

For solutions that contain non-volatile solutes, the vapor pressure of the solution can be determined by using the mole fraction of the solvent and the

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vapor pressure of the pure solvent at the same temperature..

$P_{\text{sol}} = \chi_{\text{solvent}} \cdot P_{\text{solvent}}^{\circ}$

where P_{sol} is the vapor pressure of the solution

χ_{solvent} is the mole fraction of the solvent

The vapor pressure of pure water at 25 degrees Celsius is ...

Molality is a property of

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a solution and is defined as the number of moles of solute per kilogram of solvent.

The SI unit for molality is mol/kg. A solution with a molality of 3 mol/kg is often described as “3 molal” or “3 m.” However, following the SI system of units, mol/kg or a related SI unit is now preferred.

Molality | Introduction to

Read Online Solution And Colligative **Chemistry**

This is done by cooling the liquid food below the freezing point of the solution. The freezing point depression is referred as a colligative property and it is proportional to the molar concentration of the solution (m), along with vapor pressure lowering, boiling point elevation, and osmotic pressure.

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Boiling point or freezing point of liquid solution would

...

In chemistry, an ideal solution or ideal mixture is a solution in which the gas phase exhibits thermodynamic properties analogous to those of a mixture of ideal gases. The enthalpy of mixing is zero as is the volume change on mixing by definition; the closer to

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zero the enthalpy of mixing is, the more "ideal" the behaviour of the solution becomes. The vapor pressure of the solution obeys ...

Ideal solution - Wikipedia

Osmotic Pressure Is a Property of Solutions Related to Other Colligative Properties. Osmotic pressure is closely related to some other properties of

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solutions, the colligative properties. These include the freezing point depression, the boiling point elevation, and the vapor pressure depression, all caused by dissolving solutes in a solution ...

Osmotic Pressure - an overview | ScienceDirect Topics

An isotonic solution is defined as two solutions of equal

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concentrations of solutes and water separated by a semipermeable membrane to allow water to move freely in and out of a cell.

Isotonic Solution: Definition & Example - Video & Lesson ...

Molality is a property of a solution and is defined as the number of moles of solute per kilogram of solvent.

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The SI unit for molality is mol/kg. A solution with a molality of 3 mol/kg is often described as “3 molal” or “3 m.” However, following the SI system of units, mol/kg or a related SI unit is now preferred.

Concentration Units | Chemistry [Master]

A scientific solution is defined as two or more substances in a homogenous mixture.

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Discover the parts of a solution and see examples of the three types of solutions: solid, liquid, and gas.

What is a Solution in Science? - Definition & Examples ...

Solution: Osmosis and osmotic pressure are related. Osmosis is the flow of a solvent into a solution through a semipermeable membrane. Osmotic pressure is the

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pressure that stops the process of osmosis.

Osmotic pressure is a colligative property of a substance since it depends on the concentration of the solute and not its chemical nature.

How To Calculate the Osmotic Pressure of a Solution

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